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Investigation of antioxidant and elastase inhibitory activities of *Geum urbanum* aerial parts, chemical characterization of extracts guided by chemical and biological assays

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Abstract

This study presents the bio-guided chemical investigation of the hydromethanolic (80% methanol, v/v) extract of *Geum urbanum* aerial parts. Liquid-liquid partitioning of this extract in solvents of increasing polarity combined with biological screening showed that the EtOAc soluble fraction was the most active part of the extract. This fraction was chemically profiled by a ¹³C NMR-based dereplication method, resulting in the identification of fourteen compounds. The dereplication process was followed by the purification of unknown and minor compounds of the EtOAc fraction. A new glycosylated phenol, namely 3-(3,4-dihydroxyphenyl)propyl- α -L-rhamnopyranoside, together with six known compounds were isolated. Their structures were elucidated by spectroscopic methods including NMR and HR-ESI-MS. The antioxidant activity of fractions and isolated compounds were evaluated by DPPH and hydroxyl radical scavenging, and by CUPRAC assays. In parallel, their enzyme inhibitory property against human neutrophil elastase was assessed. Four sub-fractions, essentially containing polyphenols and triterpenes, exhibited a significant elastase inhibitory activity and an ellagitannin showed a very high radical scavenging activity.

Keywords

Geum urbanum, ellagitannins, phenols, NMR dereplication, antioxidant, elastase.

Skin aging is a complex biological process mainly characterized by wrinkle formation, uneven pigmentation, darkening, thinning, and roughening of the skin. This could be caused by intrinsic (genetically determined) or extrinsic (mediated by exposure to UV radiation and to toxic compounds) mechanisms. One of the underlying mechanisms of these processes is an excessive oxidative stress, which is probably the single most harmful contributor to skin aging, leading to loss of cells and to extracellular matrix degradation as the most prominent features of chronologically aged skin. Oxidative stress results from an imbalance between reactive oxygen species production and antioxidative defense.¹ Prevention of these dynamic processes is a major issue for dermo-cosmetics and substantial efforts are being made to discover new protective ingredients.^{2,3} Plant extracts can contain high concentration of phenolic compounds, carotenoids, terpenoids, etc. able to inhibit free radicals and therefore to slow down skin aging.^{4,5} The antioxidant potency of phenolic compounds (phenolic acids, flavonoids, anthocyanidins, tannins) can be easily determined by chemical tests such as DPPH assay, the ferric reducing/antioxidant power (FRAP) assay, the ABTS assay, and the cupric ion reducing antioxidant capacity (CUPRAC) assay.⁴ Phenolics have also been reported as effective compounds in matrix remodeling activation through inhibition of elastases,⁶ a class of proteases highly involved in aging.

In the continuation of our search for local plants with anti-aging properties,^{7,8} we were interested by *Geum urbanum* L., a perennial herbaceous plant belonging to the Rosaceae family. The genus *Geum* encompasses about 70 species distributed in temperate regions all over the world. *G. urbanum* is native to humid woods of Europe and Asia⁹ and is a common species in France, especially in cool places such as hedges and undergrowth.¹⁰ Infusions from the roots and rhizomes of *G. urbanum* have been traditionally used in folk medicine in the treatment of acute diarrhea and as astringent agent for inflammations of the mucosa and gums and for treatment of hemorrhoids.^{9,11} Previous works have shown that plants species belonging to the *Geum* genus produce ellagitannins, gallotannins, flavonoids, terpenoids and phenylpropanoids.^{12,13,14} Ellagitannins (tellimagrandin, stachyurin, casuarinin, gemin A, and ellagic acid derivatives), procyanidins (procyanidin B₃, procyanidin C₂, and catechin derivatives), gallic acid and its derivatives, steroids and triterpenoids were isolated from the roots of *G. urbanum*.^{11,15,16,17,18} To our knowledge, the antioxidant and antielastase potency of *G. urbanum* aerial parts and their relation to chemical composition have not been reported. Recently, the radical scavenging activity of extracts and fractions obtained from aerial and underground parts of *G. urbanum* were investigated and the EtOAc fractions from roots and

aerial parts were found to be the most active, but only the EtOAc extract of the roots were investigated.¹⁵

Thus, the aim of the present study was to investigate the antioxidant and elastase inhibitory activities of extracts, fractions and isolated compounds obtained from the aerial parts of *G. urbanum*. A ¹³C NMR-based dereplication methodology combined to a bio-guided fractionation and purification procedure was used to identify some of its metabolites.

Results and discussion

A hydromethanolic (80% methanol, v/v) extract (hereafter referred to as GUM) was obtained from *G. urbanum* aerial parts. Its solution in water was fractionated by extraction with a series of solvents of increasing polarity, resulting in a dichloromethane fraction (DCMF), ethyl acetate fraction (EAF), and *n*-butanol fraction (*n*-BF). The free radical scavenging activities of GUM and obtained fractions were determined by the DPPH and hydroxyl radical assays whereas their cupric reducing capacity was evaluated by the CUPRAC assay. In addition, their ability to inhibit human neutrophil elastase activity was tested. The results are shown in [Table 1](#). As previously described for roots and aerial parts of *G. urbanum*,¹⁵ the EAF had the best antioxidant activities. *n*-BF exhibited the best DPPH scavenging activity (IC₅₀ < 12.5 µg/mL) followed by EAF and GUM (IC₅₀ 14.7 µg/mL and 21.8 µg/mL, respectively). The hydroxyl radicals scavenging activity was better for EAF and GUM (IC₅₀ 142.7 µg/mL and 174.5 µg/mL, respectively). For all extracts, a substantial cupric ion reducing capacity was observed (IC₅₀ < 2.0 µg/mL), but lower than the one of trolox, used as positive standard. In addition, a very good elastase inhibitory activity was observed for DCMF and *n*-BF (IC₅₀ 17.6 µg/mL and 14.3 µg/mL, respectively), but EAF showed the best activity (IC₅₀ 6.0 µg/mL), identical to quercetin used as positive standard. The EAF having interesting biological activities, in both antioxidant and anti-elastase assays, was chemically investigated through a bioassay-guided isolation strategy in order to tentatively determine the active constituents.

The major compounds of EAF were identified using a dereplication method¹⁹ combining Centrifugal Partition Chromatography (CPC), NMR analysis, clustering of NMR peak emergence profiles, and data base search, without purification of individual components. The CPC fractionation of EAF was performed with the biphasic solvent system MtBE/CH₃CN/water (3/3/4, v/v/v) to afford eleven sub-fractions (EAF₁₋₁₁). After ¹³C NMR analyses of EAF₁₋₁₁, all spectra of the fraction series were processed and submitted to Hierarchical Clustering Analysis (HCA) for the recognition of similarity between emergence

profiles of ^{13}C NMR peaks throughout the fractionation process. In this way, ^{13}C NMR signals belonging to the same compounds were grouped to build “chemical shift clusters” represented in the heat map drawn in Fig. 1. As a result, 13 major chemical shift clusters corresponding to the major metabolites of the EAF (Fig. 1), colored in yellow, were revealed by the heat map. An in-house database containing predicted chemical shift values of natural metabolites was used to identify the metabolites from the chemical shifts within clusters. Cluster 1 in fraction EAF₁₀ was assigned to saccharose and β -fructose. The identification of saccharose and β -fructose was easily confirmed by checking HSQC, HMBC and COSY data of fraction EAF₁₀ and by comparison with literature data.²⁰ Similarly, clusters 2 to 11 were identified as ellagic acid and gallic acid²¹ (clusters 2, 2' and 2"; fraction EAF₆), a mixture of 6-*O-E*-caffeoyl- β -D-glucopyranose²² and malic acid²⁰ (clusters 3 and 7; fractions EAF₂₋₈), succinic acid²³ (cluster 4; fractions EAF₃₋₄), dotorioside II²⁴ (cluster 5; fractions EAF₃₋₄), tormentic acid²⁵ (cluster 6; fraction EAF₁), caffeic acid²⁶ (cluster 7, fraction EAF₇), ellagic acid pentoside (cluster 8; fractions EAF₃₋₅), (+)-catechin (**2**)²⁷ (cluster 10; fractions EAF₂₋₃), and 3,3'-di-*O*-methylellagic acid²⁸ (cluster 11; fraction EAF₃). For the fractions EAF₄₋₅, the database proposed a 3-(3,4-dihydroxyphenyl)propyl-glycoside (**1**), which could not be identified unambiguously (cluster 9).

Fractions EAF₁₋₇ containing the phenolic compounds identified by dereplication were further screened for their antioxidant and elastase inhibitory activities. As shown in Table 1, all tested fractions presented very good antioxidant activities, revealed by DPPH and CUPRAC assays. EAF₁₋₇ showed also a significant elastase inhibitory activity, in particular EAF₁₋₄ (IC₅₀ 2.1 $\mu\text{g}/\text{mL}$, 2.9 $\mu\text{g}/\text{mL}$, 5.1 $\mu\text{g}/\text{mL}$ and 8.0 $\mu\text{g}/\text{mL}$, respectively). The most active fractions EAF₂₋₄ and EAF₇ were further purified using semi-preparative HPLC, to identify the compounds responsible of the activities, and afforded an undescribed glycosylated phenol (**1**) and six known compounds (**2** – **7**). The known compounds were identified as (+) catechin (**2**),²⁷ 1-(4-hydroxy-3-methoxyphenyl)-1,2,3-propanetriol (guaiacylglycerol) (**3**),²⁹ tellimagrandin II (**4**),³⁰ 3-*O*-methylellagic acid-4'-*O*- β -D-xylopyranoside (**5**),³¹ 2,3-dihydro-3-hydroxymethyl-7-methoxy-2-(4'-hydroxyphenyl-3'-methoxy)-5-benzofuranpropanol-4'-*O*- α -L-rhamnopyranoside (icariside E4) (**6**)³², and gemin A (**7**)³³ (Fig. 2) by comparison of physical data with literature values and from spectroscopic evidence.

Compound **1** was as yellowish-brown solid obtained in mixture with compound **3** (50/50 based in ^1H NMR). The spectral analysis of the mixture showed for **1** a molecular formula C₁₅H₂₂O₇, deduced from its negative ionization HR-ESI-MS analysis (m/z 313.1284,

[M - H]⁻). The ¹H and ¹³C NMR spectra of **1** presented resonances corresponding to aromatic and glycosidic protons and carbons. The ¹H NMR spectrum of **1** showed three signals accounting for one proton each at δ_{H} 6.65 (d, $J = 2.0$ Hz), 6.77 (d, $J = 8.2$ Hz) and 6.59 (dd, $J = 8.2, 2.0$ Hz), due to H-2', H-5' and H-6' in a tri-substituted aromatic ring. These protons were correlated in the HSQC spectrum with their aromatic carbon atoms at δ_{C} 116.4, 116.3 and 121.0, respectively. In addition, six protons at δ_{H} 3.28 (1H, m)/3.57 (1H, m), 1.76 (2H, quin, $J = 6.9$ Hz) and 2.50 (1H, q, $J = 7.3$ Hz)/2.52 (1H, q, $J = 7.2$ Hz), due respectively to the protons H-1, H-2 and H-3 of an alkyl chain correlated in the HSQC spectrum to their alkyl carbon atoms at δ_{C} 66.2, 31.1 and 31.1, respectively. Complete assignment of the remaining resonances of the aglycone in the ¹³C NMR spectrum of **1** was achieved by analysis of the HMBC data which confirmed the presence of a diphenol linked to an alkyl chain. The linkage site of the alkyl chain was determined by analysis of HMBC spectrum. The correlation between H-3 (δ_{H} 2.50/2.52) and C-1' of the aromatic ring (δ_{C} 134.7) indicated the position of alkyl chain on the diphenol. A full list of the corresponding assignments is given in experimental section. Furthermore, an anomeric proton resonance corresponding to *O*-linked sugar was observed in the ¹H NMR spectrum as doublet at δ_{H} 4.53 ($J = 1.9$ Hz, H-1'') and the signal of his corresponding ¹³C NMR carbon was at δ_{C} 100.3 (C-1''). A methyl signal at δ_{H} 1.14 (d, $J = 6.2$ Hz, H-6'') and δ_{C} 16.6 (C-6'') indicated a 6-desoxy-hexose. Analysis of COSY correlations allowed us to assign complete spin systems of an α -L-rhamnopyranose³⁴ (see experimental). The α -configuration of rhamnose deduced from the small ³ $J_{\text{H-1,H-2}}$ value was confirmed by the chemical shift of C-5 (δ_{C} 68.4).³⁵ The HMBC correlation between H-1'' (δ_{H} 4.53) and C-1 of the aglycone (δ_{C} 66.2) indicated the position of sugar on the alkyl chain. Therefore, the structure of **1** was elucidated as 3-(3,4-dihydroxyphenyl)propyl- α -L-rhamnopyranoside, as shown in Fig. 2.

Isolated compounds **2**, **5**, **6** and **7** were evaluated only for their antioxidant activity, using the DPPH scavenging assay, due to their low available mass. Compounds **1** and **3** were isolated in a mixture, so their activities could not be evaluated. The biological activity of compounds **3** and **4** was discussed based on literature data. As summarized in Table 1, compounds **2**, **5**, **6** and **7** presented a very good antioxidant activity ($\text{IC}_{50} < 23.0$ μM), with compound **7** achieving noticeably at best (IC_{50} 2.7 μM). According to literature, compounds **3**³⁶ and **4**³⁷ showed a moderate DPPH scavenging activity (**3**: IC_{50} 133.2 μM , **4**: IC_{50} 50.9 μM), while catechin (**2**) showed good antioxidant activities on DPPH (IC_{50} 18.3 μM), FRAP (IC_{50} 13.4 μM), and ABTS (IC_{50} 3.8 μM) assays.³⁸ The antioxidant activities of caffeic acid

(cluster 7), gallic acid (cluster 2) and ellagic acid are well known with IC₅₀ 58.8, 10.6 and 3.9 μM on DPPH assay, respectively.^{39,40,41} The radical scavenging activity on DPPH of dotorioside II (cluster 5) and tormentic acid (cluster 6) were also previously evaluated and showed to be very low.^{42,43} The very high activity of compound **7** is probably due to the presence of several gallic acid units containing numerous aromatic hydroxyl groups known to be related to antioxidant activity.^{40,44} Indeed, the phenols/polyphenols have the capacity to lose one or more H⁺ ions and thus to trap the free radicals in the *ortho* and *para* position of the donor hydroxyl group. In the case of a polyphenol such as catechol, the two hydroxyl groups provide two *ortho* sites and two *para* sites, which allows an efficient trapping of free radicals. The antioxidant power of compound **7** being seventeen times higher than that of compound **4**, we can suggest that the more gallic acid units in an ellagic tannins, the higher its activity.

Conclusions

The present study reports for the first time the isolation and identification of nineteen secondary metabolites including ellagic tannins, ellagic acids derivatives, phenolic acids, glycosides derivatives and triterpenes from aerial parts of *Geum urbanum*. Our studies demonstrated that EtOAc fractions (EAF) were characterized by the highest anti-elastase and antiradical activity among all prepared fractions. For this reason, we isolated individual constituents of the EAF fraction and evaluated their contribution to the observed activities. We isolated and identified a new glycosylated phenol, together with six compounds.

From a chemotaxonomic viewpoint, catechin (**2**) was isolated earlier in *G. urbanum*¹⁵ and *G. iranicum*.⁴⁵ Gemin A (**7**) was previously reported from *G. japonicum*,³³ *G. aleppicum* Jacq. and *G. calthifolium* var. *nipponicum*.⁴⁶ Among the compounds elucidated by dereplication, 6-*O-E*-caffeoyl-β-D-glucopyranose, succinic acid and 3-(3,4-dihydroxyphenyl)propyl-glycoside are reported here for the first time in the genus *Geum*. Eight compounds were previously isolated from *G. urbanum* roots: tormentic acid,¹⁵ ellagic acid pentoside,¹⁶ ellagic acid, gallic acid, caffeic acid, (+)-catechin,^{15,17} malic acid,¹⁸ and dotorioside II.¹¹ The three other compounds were already known in the genus *Geum*.^{28,45,47} Two compounds previously isolated in *G. urbanum* roots,^{15,16} the ellagic tannins **4** and **7**, could be considered as chemotaxonomic markers of this species.

Regarding biological activities, four sub-fractions, essentially containing polyphenols, presented a significant elastase inhibitory activity and, among the identified compounds,

phenolics **2**, **5**, **6** and **7** showed high antioxidant property. These tests have highlighted the antioxidant activity of polyphenols and more particularly catechin (**2**) and gemin A (**7**). The EAF and the polyphenols identified from *G. urbanum* are promising compounds for applications in the dermo-cosmetic field due to their very interesting antioxidant and anti-elastase activities.

Experimental

General experimental procedures

1D- and 2D-NMR spectra were recorded at 298 K in CH₃OH-*d*₄ or DMSO-*d*₆ on a Bruker AvanceIII-600 spectrometer (Karlsruhe, Germany) (¹H at 600 MHz and ¹³C at 151 MHz) equipped with a 5 mm TCI cryoprobe. The Bruker TopSpin 3.6 software was used for NMR data acquisition and processing. HR-ESI-MS spectra were recorded on a Micromass Q-TOF micro instrument (Manchester, UK). Mass spectra were recorded in the negative ionization mode in the *m/z* range 100–2000, with a mass resolution of 20.000 and an acceleration voltage of 0.7 kV. Semi-preparative HPLC purification was carried out on a Dionex apparatus equipped with an ASI-100 automated sample injector, a STH 585 column oven, a P580 pump, and a UVD 340S diode array detector, all driven by the Chromeleon® software version 6.8. Analytical HPLC experiments were performed using a ThermoFisher Ultimate 3000 (Thermo Fischer Scientific, Villebon-sur-Yvette, France), equipped with a four ways pump LPG 3400 SD, an automatic injector WPS 3000 SL, a UV/visible diode array detector 3000, and the Chromeleon® software version 6.8. A prepacked C₁₈ column (Interchim, 250 × 10 mm, 5 μ) was employed for semi-preparative HPLC. The mobile phase was composed of H₂O acidified with TFA (0.025%) and CH₃CN with a flow rate of 5 mL/min and the chromatograms were monitored at 205, 254, 300 and 360 nm. A prepacked C₁₈ column Uptisphere Strategy C₁₈ (Interchim, 250 × 4.6 mm, 5 μ) was used for analytical HPLC and the mobile phase contained H₂O with TFA (0.025% v/v) and CH₃CN. A gradient elution method was applied, from 5% to 80% of CH₃CN in 30 min with a flow rate of 1 mL/min and the chromatograms were monitored at 205, 254, 300 and 360 nm. Thin-layer chromatography (TLC) was carried out on silica gel 60 F₂₅₄ pre-coated aluminum plates (0.2 mm, Merck), using CHCl₃/MeOH/H₂O (14/6/1, v/v/v) as mobile phase. The compounds were visualized under UV light (254 and 366 nm) after high-performance thin-layer chromatography (HPTLC, CAMAG TLC Visualizer 2) and sprayed with 50% H₂SO₄ followed by heating. A FLUOstar Omega spectrophotometer (BMG LABTECH) was used for absorbance measurement in antioxidant assays. An Infinite

F200 PRO spectrofluorimeter (Tecan, Lyon, France) was used for measuring fluorescence during anti-elastase assays.

Plant material

The aerial parts of *Geum urbanum* L. (Rosaceae) were collected in Châlons-sur-Vesle (Northeastern part of France: 49°21'67"N, 4°05'E) in July 2016, authenticated by Dr. Abdulmagid Alabdul Magid and dried at room temperature. A voucher specimen (MA-GU-2016-07) was deposited at the Herbarium of the Botanic Laboratory-Faculty of Pharmacy, University of Reims Champagne-Ardenne.

Extraction and isolation

The dried and powdered *G. urbanum* aerial parts (300 g) were macerated in MeOH/water (4/1, v/v, 3 × 3 L, 24 h) at room temperature. The macerate was concentrated to about 1 L at 40 °C under vacuum. An aliquot of the aqueous solution (100 mL) was evaporated to dryness to obtain the GUM extract (2.6 g). The aqueous solution (1 L) of GUM was successively extracted with dichloromethane (3 × 750 mL), ethyl acetate (3 × 750 mL) and *n*-butanol (3 × 750 mL), then dried under reduced pressure to yield DCMF (6.1 g), EAF (6.2 g) and *n*-BF (24.2 g) extracts, respectively and a water-soluble part (29.2 g).

Centrifugal partition chromatography

Centrifugal partition chromatography (CPC) fractionation was carried out on a lab-scale FCPE300® column of 303 mL inner volume (Rousselet Robatel Kromaton, Annonay, France) made of 7 circular partition disks and engraved with a total of 231 partition twin-cells (\approx 1 mL per twin cell). The liquid phases were pumped by a KNAUER Preparative 1800 V7115 pump (Berlin, Germany). The column was coupled on-line with a UVD 170 S detector set at 210, 254, 280 and 366 nm (Dionex, Sunnivale, CA, USA). Fractions of 20 mL were collected by a Pharmacia Superfrac collector (Uppsala, Sweden). The solvent system was MtBE/CH₃CN/water (3/3/4, v/v/v). The column rotation speed was set at 1200 rpm and the flow rate at 20 mL/min. EAF (1.5 g injected mass) was dissolved in 15 mL of a mixture of both lower phase (10 mL) and upper phase (5 mL). The upper phase of the biphasic solvent system was pumped for 85 min in the ascending mode. The stationary phase was then extruded by pumping the organic phase in the descending mode at 20 mL/min. Fractions of 20

mL were collected over the whole process. All fractions were analyzed by TLC and HPLC and then pooled to yield fractions EAF₁₋₁₁.

NMR analysis and dereplication of the metabolites

The structures and names of metabolites already reported in the genus *Geum* (n=50) were collected from articles available in the scientific literature. The predicted ¹³C NMR chemical shifts of each one was then stored into a local database already comprising 3038 structures of natural compounds (NMR Workbook Suite 2012, ACD/Labs, Ontario, Canada). In a second step, all CPC fractions were dried under vacuum and an aliquot (up to 20 mg when possible) was dissolved in 600 μL of DMSO-*d*₆ and analyzed by ¹³C NMR. A standard zgpg pulse sequence was used with an acquisition time of 0.9 s and a relaxation delay of 3 s. For each sample, 1024 scans were added to obtain a satisfactory signal-to-noise ratio. The spectral width was 240 ppm and the receiver gain was set to the highest possible value. Spectra were then manually phased, baseline corrected, and calibrated on the central resonance of DMSO-*d*₆ (δ 39.8 ppm). The last step consisted in the binning of all ¹³C NMR signals followed by the visualization of the whole dataset as a heat map. For this purpose, the absolute intensities of all ¹³C NMR signals detected in the spectra of the fraction series were automatically collected and each resulting peak list was stored as a text file. The binning step was performed by a locally developed computer script written in Python language. Its principle was to divide the ¹³C spectral domain (from 0 to 240 ppm) into chemical shift intervals (or bins) of identical width ($\Delta\delta = 0.2$ ppm), and to associate the absolute intensity of each peak to the corresponding bin. The resulting table was imported into the PermutMatrix version 1.9.3 software (LIRMM, Montpellier, France) and submitted to Hierarchical Clustering Analysis (HCA) for data visualization. The chemical shift clusters regrouped with the HCA were then submitted to database search for compound identification. Additional 1D and 2D NMR experiments (¹H NMR, HSQC, HMBC, and ¹H-¹H-COSY) were recorded and analyzed in order to confirm the structures of the identified compounds.

HPLC purification of EAF

Fraction EAF₂ was subjected to semi-preparative HPLC using the gradient system (15-40% CH₃CN, 30 min) to yield compounds **2** (*t*_R=5.9 min, 3.8 mg), **4** (*t*_R=9.5 min, 4.3 mg) and the mixture of **1** and **3** (*t*_R=20.2 min, 5.1 mg). Fractions EAF₃₋₄ were purified by semi-prep. HPLC with gradient system (10-40% CH₃CN, 30 min) to yield compounds **1** and **3** in mixture

(t_R =12.6 min, 2.7 mg), **5** (t_R =13.4 min, 3.8 mg) and **6** (t_R =13.8 min, 3.7 mg). Fraction EAF₇ was purified by semi-prep. HPLC with gradient system (10-25% CH₃CN, 20 min) to yield compound **7** (t_R =11.2 min, 4 mg).

4.4. 3-(3,4-Dihydroxyphenyl)propyl- α -L-rhamnopyranoside (**1**)

Yellowish-brown solid; ¹H NMR (600 MHz, CD₃OD, supplemental data, figure S1) 3.28 (1H, m, H-1a), 3.57 (1H, m, H-1b), 1.76 (1H, quin, J =6.9 Hz, H-2a), 2.50 (1H, qd, J =7.3 Hz, H-3a), 2.52 (1H, qd, J =7.2 Hz, H-3b), 6.65 (1H, d, J =2.0 Hz, H-2'), 6.77 (1H, d, J =8.2 Hz, H-5'), 6.59 (1H, dd, J =8.2,2.0 Hz, H-6'), 4.53 (1H, d, J =1.9 Hz, H-1''), 3.69 (1H, dd, J =3.4,1.9 Hz, H-2''), 3.55 (1H, dd, J =9.4,3.4 Hz, H-3''), 3.26 (1H, t, J =9.4 Hz, H-4''), 3.48 (1H, m, H-5''), 1.14 (3H, d, J =6.2 Hz, H-6''); ¹³C NMR (125 MHz, CD₃OD, supplemental data, figure S2) 66.2 (C-1), 31.1 (2), 31.2 (3), 134.7 (1'), 116.4 (C-2'), 143.7 (C-3'), 141.6 (C-4'), 116.3 (C-5'), 121.0 (C-6'), 100.3 (C-1''),70.9 (C-2''), 71.0 (C-3''), 72.6 (C-4''), 68.4 (C-5''), 16.6 (C-6''); HR-ESI-MS (negative-ion mode) m/z : 313.1284 [M - H]⁻ (calculated for C₁₅H₂₁O₇, 313.1287).

DPPH radical scavenging activity

Extracts, fractions and compounds **2**, **5**, **6** and **7** were evaluated for their DPPH radical scavenging activity according to a recently published procedure.⁸ Samples were prepared at concentrations of 200, 100, 50, 25, 12.5 and 6.25 μ g/mL. Ascorbic acid and quercetin were used as positive controls. All the tests were conducted in triplicate for each concentration.

Hydroxyl radical scavenging activity

Extracts and fractions were evaluated for their hydroxyl radical scavenging activity at various concentrations, ranging from 1330 to 41.56 μ g/mL, according to a recently published procedure.⁸

Power cupric ion reducing (CUPRAC) assay

Extracts and fractions were tested for their cupric ion reducing ability. The cupric ion reducing activity (CUPRAC) was determined according to a published method.⁴⁸ Samples were prepared at concentrations of 572, 286, 143, 71.5, 35.75, 17.87 and 8.94 μ g/mL and dissolved in H₂O/DMSO (9/1, v/v). A volume of 45 μ L of each diluted sample was added to a premixed reaction mixture containing CuCl₂ (90 μ L, 10 mM), freshly prepared neocuproine (90 μ L, 7.5 mM, dissolved in distilled water and ethanol in proportion 8/2, (v/v) and NH₄Ac

buffer (90 μL , 1M, pH 7.0). Similarly, a blank test mixture was prepared by adding sample solution (45 μL) to premixed reaction mixture (270 μL) without CuCl_2 . The reaction proceeded for 30 min at room temperature on a 96-well microplate and the absorbance was then read at 450 nm. Ascorbic acid, quercetin and trolox were used as positive controls. The power cupric ion reducing was calculated as follows: $[1 - A_0 / (A_1 - A_2)] \times 100$, where A_0 is absorbance of the control (without sample), A_1 is absorbance in the presence of the sample and A_2 is absorbance of the blank. All the tests were conducted in triplicate and IC_{50} were determined by interpolation of concentration versus inhibition curves obtained by a MSeExcel calculation sheet.

Elastase enzyme assay

Extracts and fractions were evaluated for their ability to inhibit Human Leukocyte Elastase (HLE) (Merck Biosciences). Tests were performed in 96-well microplates pre-coated with 1% Bovine Serum Albumin. HLE (0.8 μM) was incubated for 1 h at 27°C in Tris buffer (50 mM Tris-HCl pH 7.5 containing 500 mM NaCl) containing 0.1 to 1000 $\mu\text{g/mL}$ of tested sample. Pure solvent was used as a control sample. The assay was initiated by adding HLE fluorogenic substrate MeOSuc-Ala-Ala-Pro-Val-AMC ($\lambda_{\text{exc}} = 380 \text{ nm}$ / $\lambda_{\text{em}} = 460 \text{ nm}$) at a final concentration of 80 μM . The rate of substrate cleavage was measured in triplicate for each sample concentration with one measurement per minute for 60 min. HLE activity was calculated according to: % HLE activity = $(\text{Slope}_{\text{sample}} \times 100) / \text{Slope}_{\text{control}}$, where $\text{slope}_{\text{sample}}$ and $\text{slope}_{\text{control}}$ are the slope of the curve in the graph of fluorescence intensity versus time. The IC_{50} values were calculated by non-linear regression analysis with the Graphpad software (La Jolla, USA).

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Declaration of Conflicting Interests

The authors declare no conflict of interest.

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Supplemental Material

Supplemental material for this article is available online.

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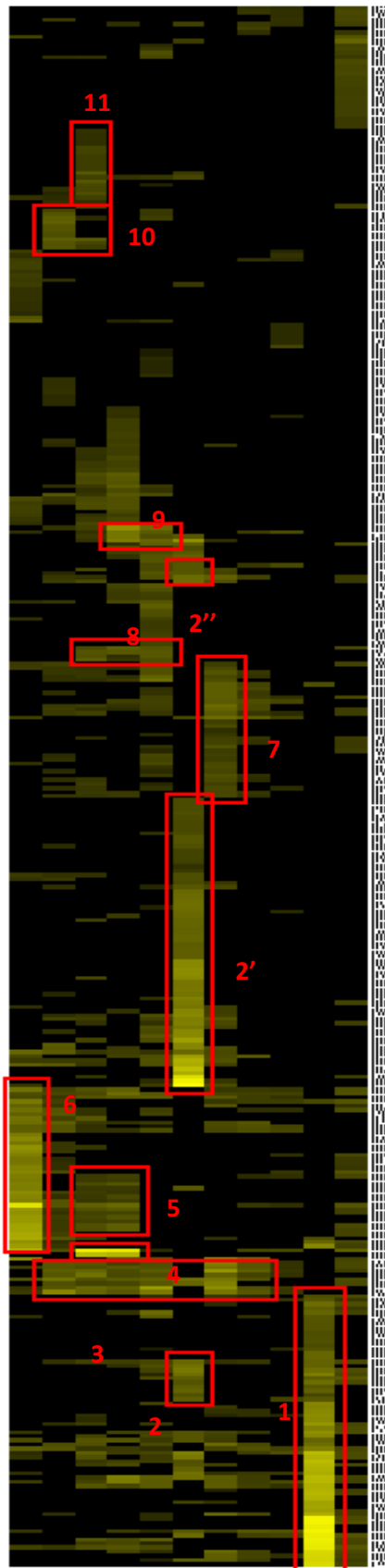
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Table 1

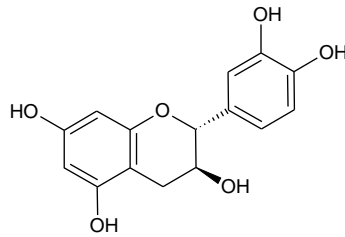
Antioxidant and elastase inhibitory activities of extracts, fractions and compounds isolated from *G. urbanum*.

	DPPH radical scavenging activity		OH radical scavenging activity IC ₅₀ (µg/mL)	Power cupric ion reducing (CUPRAC) IC ₅₀ (µg/mL)	Human neutrophil elastase inhibition IC ₅₀ (µg/mL)
	IC ₅₀ (µg/mL)	IC ₅₀ (µM)			
GUM	21.8 ± 1.3		174.5 ± 1.8	1.5 ± 0.1	110.4 ± 1.1
DCMF	(42%) ^a		188.3 ± 3.5	1.8 ± 0	17.6 ± 1.1
EAF	14.7 ± 0.8		142.7 ± 5.7	0.2 ± 0	6.0 ± 1.1
n-BF	< 12.5		263.0 ± 3.3	0.5 ± 0	14.3 ± 1.1
EAF₁	13.8 ± 0.1		365.0 ± 3.1	4.8 ± 0.3	2.1 ± 1.2
EAF₂	8.9 ± 0.6		199.2 ± 3.8	2.9 ± 0.1	2.9 ± 1.1
EAF₃	11.5 ± 0		309.0 ± 4.8	2.3 ± 0	5.1 ± 1.1
EAF₄	3.4 ± 0.1		139.0 ± 4.6	3.0 ± 0	8.0 ± 1.1
EAF₅	3.5 ± 0.1		129.5 ± 3.3	1.8 ± 0	13.1 ± 1.1
EAF₆	3.0 ± 0.1		131.5 ± 0	1.2 ± 0	11.9 ± 1.1
EAF₇					12.2 ± 1.2
2	22.6 ± 2.2	77.9 ± 7.6			
3	28.5 ± 0.8 ³⁶	133.1 ± 3.7 ³⁶			
4	47.8 ± 3.1 ³⁷	50.9 ± 3.3 ³⁷			
5	20.1 ± 0.4	44.9 ± 0.9			
6	18.5 ± 0.2	36.5 ± 0.4			
7	5.0 ± 0.3	2.7 ± 0.2			
Quercetin^b	5.4 ± 0.2	17.9 ± 0.7	52.2 ± 4.8	13.6 ± 1.3	6.0 ± 0 ⁴⁹
Ascorbic acid^b	2.3 ± 0.4	13.1 ± 2.3	229.2 ± 2.3	13.3 ± 0.5	
Trolox^b				5.4 ± 0.3	

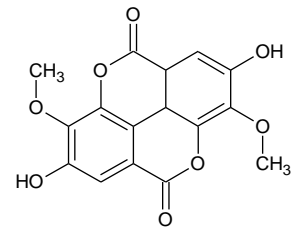
a % Inhibition at 200 µg/mL, b used as positive control.



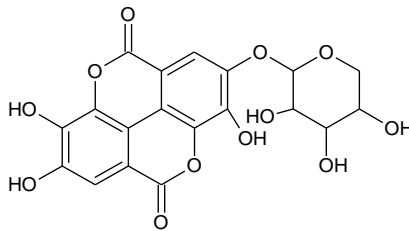
Cluster 10: (+)-Catechine



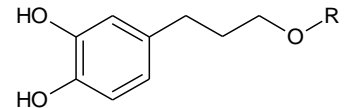
Cluster 11: 3,3'-di-O-methylellagic acid



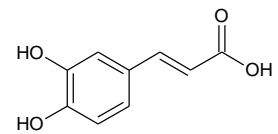
Cluster 8 : Ellagic acid pentoside



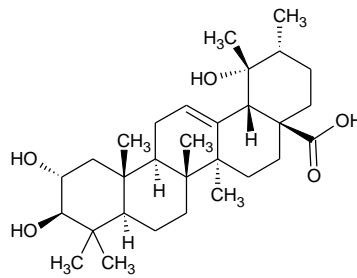
Cluster 9: 3-(3,4-Dihydroxyphényl) propyl-glycosylated



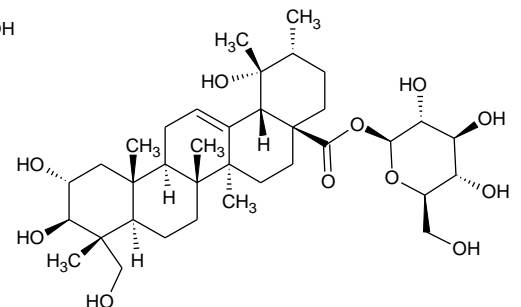
Cluster 7: Caffeic acid



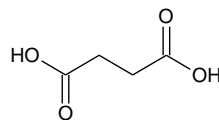
Cluster 6: Tormentic acid



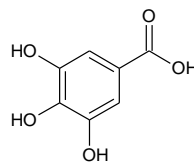
Cluster 5 : Dotorioside II



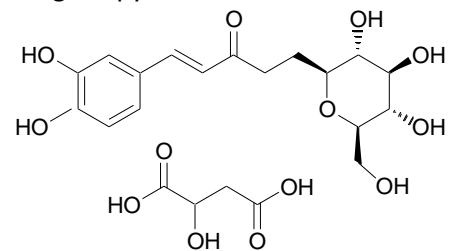
Cluster 4 : Succinic acid



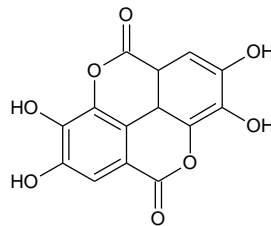
Cluster 2': Gallic acid



Cluster 3 + 7: 6-O-trans-Caffeoyl-β-D-glucopyranoside + malic acid



Cluster 2 + 2' + 2'' : Ellagic acid



Cluster 1: Saccharose + β-D-Fructose

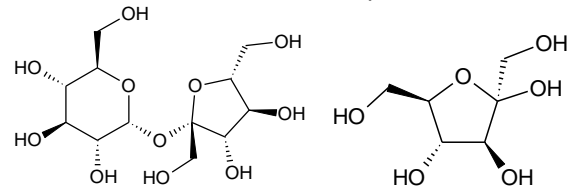


Fig. 1. ¹³C NMR chemical shift clusters obtained by applying HCA on EAF CPC fractions

from *G. urbanum*.

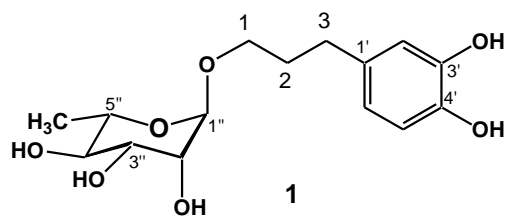


Fig. 2. Chemical structure of compound **1** isolated from *G. urbanum* aerial parts.